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S(s) + O2(g) ----> SO2(g) Pass this with more air over a vanadium(V) oxide catalyst at a temperature of 450°C and a pressure of 1 to 2 atmospheres to give sulfur trioxide. 2SO2(g) + O2(g) 2SO3(g) Absorb the sulfur trioxide in concentrated sulfuric acid to give fuming sulfuric acid.

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11 x N = 14; 3 x H = 3; 1 x S = 32; 3 x O = 16 x 3 = 48. 2Empirical mass =14 + 3 + 32 + 48 = 97. 3Divide molar mass by empirical mass: 97.01/97 = 1, therefore molecular formula = empirical formula. 42© Pearson Education Ltd 2015. Copying permitted for purchasing institution only. This material is not copyright free.

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